

## Molecular Compounds Practice Problems

I. Write names for each of the following molecular substances.

1.  $\text{SiO}_2$  : \_\_\_\_\_

2.  $\text{PCl}_3$  : \_\_\_\_\_

3.  $\text{SiF}_4$  : \_\_\_\_\_

4.  $\text{N}_2\text{O}$  : \_\_\_\_\_

5.  $\text{SO}_3$  : \_\_\_\_\_

6.  $\text{N}_2\text{O}_5$  : \_\_\_\_\_

7.  $\text{IF}_5$  : \_\_\_\_\_

8.  $\text{SF}_6$  : \_\_\_\_\_

9.  $\text{ClO}_2$  : \_\_\_\_\_

10.  $\text{P}_4\text{S}_3$  : \_\_\_\_\_

11.  $\text{NO}$  : \_\_\_\_\_

12.  $\text{SF}_4$  : \_\_\_\_\_

13.  $\text{XeF}_4$  : \_\_\_\_\_

14.  $\text{SbF}_5$  : \_\_\_\_\_

15.  $\text{NH}_3$  : \_\_\_\_\_

16.  $\text{SO}_2$  : \_\_\_\_\_

17.  $\text{H}_2\text{O}$  : \_\_\_\_\_

18.  $\text{CS}_2$  : \_\_\_\_\_

19.  $\text{Cl}_4$  : \_\_\_\_\_

20.  $\text{BCl}_3$  : \_\_\_\_\_

**II. Write the formula for the following molecular compounds.**

1. dichlorine monoxide : \_\_\_\_\_

2. chlorine dioxide : \_\_\_\_\_

3. carbon disulfide : \_\_\_\_\_

4. selenium dichloride : \_\_\_\_\_

5. chlorine trifluoride : \_\_\_\_\_

6. diboron tetroxide : \_\_\_\_\_

7. sulfur hexafluoride : \_\_\_\_\_

8. diphosphorous trisulfide : \_\_\_\_\_

9. phosphorus tetrachloride : \_\_\_\_\_

10. silicon dioxide : \_\_\_\_\_

11. nitrogen dioxide : \_\_\_\_\_

12. sulfur hexaiodide : \_\_\_\_\_

13. carbon tetrachloride : \_\_\_\_\_

14. nitrogen triiodide : \_\_\_\_\_

15. sulfur trioxide : \_\_\_\_\_

16. phosphorus pentachloride : \_\_\_\_\_

17. nitrogen trifluoride : \_\_\_\_\_

18. carbon tetrafluoride : \_\_\_\_\_

19. sulfur hexafluoride: \_\_\_\_\_

20. dinitrogen tetroxide : \_\_\_\_\_