

## Naming Compounds Practice Problems

### I. Identify the names of the following ionic compounds.

1. NaCl \_\_\_\_\_
2. MgCl<sub>2</sub> \_\_\_\_\_
3. LiOH \_\_\_\_\_
4. H<sub>2</sub>O<sub>2</sub> \_\_\_\_\_
5. Na<sub>2</sub>S \_\_\_\_\_
6. Al<sub>2</sub>O<sub>3</sub> \_\_\_\_\_
7. SrBr<sub>2</sub> \_\_\_\_\_
8. Cu<sub>2</sub>(SO<sub>4</sub>) \_\_\_\_\_
9. Cr(C<sub>2</sub>H<sub>3</sub>O<sub>2</sub>)<sub>2</sub> \_\_\_\_\_
10. (NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub> \_\_\_\_\_

### II. Write the formulas for the following ionic compounds.

1. magnesium oxide \_\_\_\_\_
2. potassium iodide \_\_\_\_\_
3. sodium sulfide \_\_\_\_\_
4. tin (II) fluoride \_\_\_\_\_
5. aluminum chloride \_\_\_\_\_
6. iron (III) bromide \_\_\_\_\_
7. calcium sulfate \_\_\_\_\_
8. sodium hydroxide \_\_\_\_\_
9. potassium permanganate \_\_\_\_\_
10. potassium dichromate \_\_\_\_\_

**III. Identify the names of the following covalent molecules.**

1.  $\text{CO}_2$  \_\_\_\_\_
2.  $\text{H}_2\text{O}$  \_\_\_\_\_
3.  $\text{CO}$  \_\_\_\_\_
4.  $\text{PCl}_3$  \_\_\_\_\_
5.  $\text{P}_4\text{O}_6$  \_\_\_\_\_
6.  $\text{SiO}_2$  \_\_\_\_\_
7.  $\text{CCl}_4$  \_\_\_\_\_
8.  $\text{N}_2\text{F}_2$  \_\_\_\_\_
9.  $\text{XeF}_2$  \_\_\_\_\_
10.  $\text{Cl}_2\text{O}_7$  \_\_\_\_\_

**IV. Write the formulas for the following molecular compounds.**

1. nitrogen tribromide \_\_\_\_\_
2. sulfur dioxide \_\_\_\_\_
3. dichlorine monoxide \_\_\_\_\_
4. dinitrogen tetrafluoride \_\_\_\_\_
5. sulfur hexafluoride \_\_\_\_\_
6. dihydrogen sulfide \_\_\_\_\_
7. dinitrogen pentoxide \_\_\_\_\_
8. carbon tetrabromide \_\_\_\_\_
9. diarsenic pentasulfide \_\_\_\_\_
10. tetraphosphorus decasulfide \_\_\_\_\_