

Types Of Reactions Practice Problems

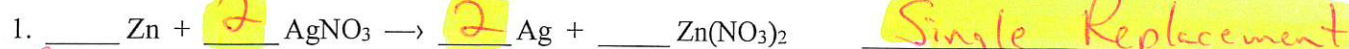
I. Balance the equation.

* = Combustion Reaction

II. Determine whether each reaction is : Synthesis, Decomposition, Single Replacement or Double Replacement.

III. Write the names the elements and compounds involved with each reaction.

Type Of Reaction



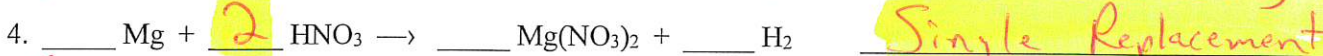
(zinc + silver nitrate → silver + zinc nitrate)



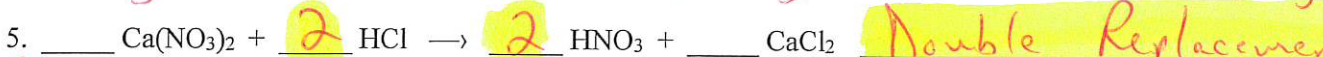
(iron (III) hydroxide → iron (III) oxide + water)



(ammonia + sulfuric acid → ammonium sulfate)



(magnesium + nitric acid → magnesium nitrate + hydrogen)



(calcium nitrate + hydrochloric acid → nitric acid + calcium chloride)



(sodium + water → sodium hydroxide + hydrogen)



(carbon monoxide + oxygen → carbon dioxide)



(iron (II) sulfide + hydrochloric acid → iron (II) chloride + hydrogen sulfide)



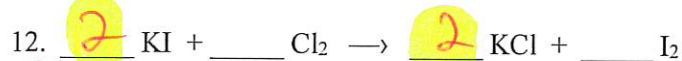
(sodium nitrate → sodium nitrite + oxygen)



(methane + oxygen → carbon dioxide + water)



(iron + copper (I) nitrate → copper + iron (II) nitrate)



(potassium iodide + chlorine → potassium chloride + iodine)

Single Replacement



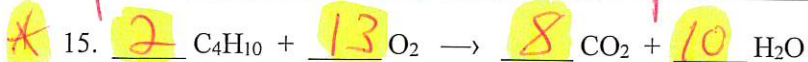
(aluminum + sulfur → aluminum sulfide)

Synthesis



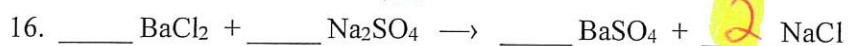
(potassium chlorate → potassium chloride + oxygen)

Decomposition



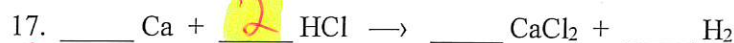
(butane + oxygen → carbon dioxide + water)

Decomposition/Synthesis



(barium chloride + sodium sulfate → barium sulfate + sodium chloride)

Double Replacement



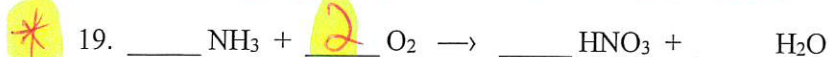
(calcium + hydrochloric acid → calcium chloride + hydrogen)

Single Replacement



(zinc chloride + ammonium sulfide → zinc sulfide + ammonium chloride)

Double Replacement



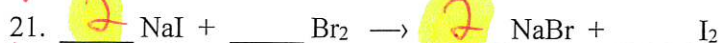
(ammonia + oxygen → nitric acid + water)

Decomposition/Synthesis



(potassium + water → potassium hydroxide + hydrogen)

Single Replacement



(sodium iodide + bromine → sodium bromide + iodine)

Single Replacement



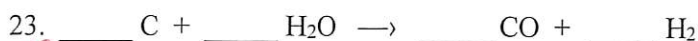
(silver + sulfur → silver sulfide)

Synthesis



(sodium chlorate → sodium chloride + oxygen)

Decomposition



(carbon + water → carbon dioxide + hydrogen)

Single Replacement

